

Kc Calculations 1 Chemsheets

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The equilibrium below was established by allowing 200 moles of ethanol and 100 moles of ethanoic acid to react at 25°C. At equilibrium the mixture contained 0.845 moles of ethyl ethanoate. Calculate K_c at this temperature.

$$\text{CH}_3\text{COOH}(\text{l}) + \text{CH}_3\text{CH}_2\text{OH}(\text{l}) \rightleftharpoons \text{CH}_3\text{COOCH}_2\text{CH}_3(\text{l}) + \text{H}_2\text{O}(\text{l})$$

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Chemsheets AS 1044 c) State Le Châteliers

C h e m g u i d e - q u e s t i o n s

C h e m g u i d e - q u e s t i o n s K C 1 The reaction between ethene and steam to produce ethanol is an example of a homogeneous equilibrium:

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© www.CHEMSHEETS.co.uk 14-Jul-12 Chemsheets A2 009 7 TASK 3 - pH of strong bases 1) Calculate the pH of the following solutions a) 0.15 mol dm⁻³ KOH b) 0.05 mol

K PROBLEMS - Chemsheets

1) 1000 moles of PCl₅ vapour are heated to 500 K in a sealed vessel. The equilibrium mixture, at a pressure of 625 kPa, contains 0.600 moles of chlorine. Calculate K_p.

$$\text{PCl}_5(\text{g}) \rightleftharpoons \text{PCl}_3(\text{g}) + \text{Cl}_2(\text{g})$$

2) A mixture of 190 moles of H₂ and 190 moles of I₂ were allowed to reach equilibrium at 710 K. The equilibrium mixture

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© www.CHEMSHEETS.co.uk 21-Jan-2018 Chemsheets A2 1001 Page 2 SECTION 1 - Recap of AS Kinetics What is reaction rate? The rate of a

chemical reaction is a measure

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Full worked solutions are available to subscribers of www.chemsheets.co.uk Subscribe for many more exercises with answers TASK 1 - Bronsted-Lowry acids & bases

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• The amount of energy needed to make 1 g of a substance 1 °C (1 K) hotter is called the specific heat capacity (measured in J g⁻¹ K⁻¹) • The following equation is then used to find the amount of heat energy give out (or absorbed)

Chemsheets A2 033 Thermodynamics

1 k-1 8 25 cm³ of 20 mol dm⁻³ nitric acid was reacted with 25 cm³ of 20 mol dm⁻³ potassium hydroxide in an insulated cup The temperature rose from 20°C to 33.9°C

Chemsheets AS 029 (Thermodynamics) ANS 2

TASK 1 - Calorimetry calculations 1 10 -1984 kJ mol⁻¹ Full worked solutions in: Chemsheets AS029 (Calorimetry) TASK 2 - Calorimetry calculations 2 1 -3800 kJ mol⁻¹ 2 -518 kJ mol⁻¹ 3 -2500 kJ mol⁻¹ Microsoft Word - Chemsheets AS 029 (Thermodynamics) ANS 2

5.9 10 -3 - Welcome to web.gccaz.edu

CHM152 Equilibrium Worksheet Key 4 Note: Different initial amounts and large K so short cuts don't apply! [H₂] = L moles 3 000 3 000

3. equilibria - chemrevise

Apr 03, 2014 · If the K_c has no unit then there are equal numbers of reactants and products In this case you do not have to divide by volume to work out concentration and equilibrium moles could be put straight into the K_c expression $K_c = \frac{[\text{HCl}(\text{g})]^2}{[\text{H}_2(\text{g})][\text{Cl}_2(\text{g})]}$ N Goalby chemrevise.org 2

Initially there were 15 moles of N₂ and 4 mole of H₂, in a 1

WORKSHEET: CHEMICAL EQUILIBRIUM Name Last First

3 For the equilibrium: $2 \text{BrCl}(\text{g}) \rightleftharpoons \text{Br}_2(\text{g}) + \text{Cl}_2(\text{g})$ at 205 °C, the equilibrium constant, K_c, is 0.143 If 134 moles each of Br₂(g) and Cl₂(g) are introduced in a container which has a volume of 110 liters and allowed to reach equilibrium at 205 °C, what would be

Worksheet 2-3 Calculations Involving the Equilibrium ...

Worksheet 2-3 - Calculations Involving the Equilibrium Constant Page 9 15 The equilibrium constant for the reaction: $\text{H}_2(\text{g}) + \text{I}_2(\text{g}) \rightleftharpoons 2\text{HI}(\text{g})$ at 448°C is 50 a) If 10 mol of H₂ is mixed with 10 mol of I₂ in a 0.50 L container and allowed to react at 448°C, what is the equilibrium [HI]? Answer _____

115 Lab EQUILIBRIUM WORKSHEET Name: a.) 2 N O (g) 4 NO ...

Page 1 of 6 115 Lab EQUILIBRIUM WORKSHEET Name: _____ 1 Write the law of mass action for the following reactions: a) $2 \text{N}_2\text{O}_5(\text{g}) \rightleftharpoons 4 \text{NO}_2(\text{g}) + \text{O}_2(\text{g})$ with K_c = 100 x 10⁻⁶ What would be the equilibrium composition at the temperature for which the above K_c is valid? 16

Chem 1711 Born-Haber Cycle, Practice Problems

Chem 1711 Born-Haber Cycle, Practice Problems 1 Given the following information for magnesium, oxygen, and magnesium oxide calculate the second electron gain

Chem guide - answers BOND ENTHALPIES

Chem guide - answers BOND ENTHALPIES 1 a) It means that it takes 193 kJ to break the bonds in 1 mole of gaseous bromine to produce gaseous

Practical Guide AQA

volume of 1 drop of solution delivered from a burette and so this is the smallest difference in readings that can be measured If the bottom of the meniscus sits on a line it should end with a 000 as in the above example 900cm³

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